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**SOLVENT EXTRACTION BATCH DISTRIBUTION
COEFFICIENTS WITH SAVANNAH RIVER SITE DISSOLVED
SALT CAKE**

**D. D. Walker
M. H. Beasley
A. D. Coleman
B. H. Croy
C. C. DiPrete
D. P. DiPrete**

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**Westinghouse Savannah River Company
Savannah River Technology Center
Aiken, SC 29808**



SAVANNAH RIVER SITE

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REVIEWS AND APPROVALS

Authors

D. D. Walker Date
Waste Processing Technology

M. H. Beasley Date

A. D. Coleman Date

B. H. Croy Date
Waste Processing Technology

C. C. DiPrete Date
Analytical Development

D.P. DiPrete Date
Analytical Development

Technical Review

W. R. Wilmarth Date
Waste Processing Technology

Approvals

M. C. Thompson, SRTC, TFA System Lead Date

R. E. Edwards, Manager, HLW Process Engineering Date

J. T. Carter, Director of Engineering, Salt Processing Project Date

H. D. Harmon, TFA SPP Technology Development Manager Date

W. D. Clark, DOE-SR, Technical Support Manager Date

S. D. Fink, Manager, Liquid Waste Processing Group Date

W. L. Tamosaitis, Manager, Waste Processing Technology Section Date

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D. D. Walker, M. H. Beasley, A. D. Coleman, B. H. Croy, C. C. DiPrete,
and D. P. DiPrete

SUMMARY

Researchers characterized high-level waste derived from dissolved salt cake from the Savannah River Site (SRS) tank farm and measured the cesium distribution coefficients (D_{Cs}) for extraction, scrub, and stripping steps of the caustic-side solvent extraction (CSSX) flowsheet. The measurements used two SRS high-level waste samples derived entirely or in part from salt cake. The chemical compositions of both samples are reported. Dissolved salt cake waste contained less Cs-137 and more dianions than is typical of supernate samples. Extraction and scrub D_{Cs} values for both samples exceeded process requirements and agreed well with model predictions. Strip D_{Cs} values for the Tank 46F sample also met process requirements. However, strip D_{Cs} values could not be calculated for the Tank 38H sample due to the poor material balance for Cs-137. Potential explanations for the poor material balance are discussed and additional work to determine the cause is described.

INTRODUCTION

A solvent extraction process for removal of cesium from alkaline solutions has been developed utilizing a novel solvent invented at the Oak Ridge National Laboratory (ORNL).¹ This solvent consists of a calix[4]arene-crown-6 extractant dissolved in an inert hydrocarbon matrix. A modifier, an alkylphenoxy alcohol, is added to the solvent to enhance the extraction power of the calixarene and to prevent the formation of a third phase. An additional additive, trioctylamine, is used to improve stripping performance and to mitigate the effects of any surfactants present in the feed stream.

A number of laboratory studies with simulated wastes determined that the solvent system is stable and performs well in the presence of the major chemical components in the SRS high-level waste tanks. However, confirmation of the simulant results with actual high-level waste samples from the SRS tank farms have yielded mixed results relative to cesium removal efficiency. In general the extraction D_{Cs} values equal or exceed the flowsheet requirement of 8.²⁻³ Nevertheless, the experimental values tend to be lower than the values predicted by a model based on waste simulant measurements developed by ORNL researchers.⁴ Previous results are summarized in Table I.

TABLE I. D_{Cs} Values with Actual SRS High-Level Waste*

Tank	D _{Cs} Values (25 °C)						Ref.
	Extraction		Scrub	Strip			
	Model**	Expt.		1	2	3	
13H	8.8	10.1	4.4	1.1	0.25	0.12	3
26F	15	7.7	1.1	0.10	0.23	0.04	3
33F	24	8.8	1.1	0.07	0.04	0.06	3
35H	22	7.9	0.7	0.14	0.17	0.16	3
46F	17	13.8	1.6	0.27	0.10	0.04	3
37F/44H		19.0	1.76	0.20	0.10	0.07	2

* Supernate samples.

**ORNL model, Ref. 4

This work partially fulfills a request from the U. S. Department of Energy, Savannah River Operations Office (C. E. Anderson, Assistant Manager for High-Level Waste, Department of Energy-Savannah River, memorandum to S. Wood, "Fiscal Year (FY) 2001 Annual Operating Plan (AOP) Changes in Support of the Salt Processing Project," dated 10 July 2001). The work complies with the following task plan: D. D. Walker, "Task Technical and Quality Assurance Plan for CSSX Real Waste Batch Tests," WSRC-RP-2001-00772, Rev.0, August 14, 2001.

EXPERIMENTAL

Tank 38H and 46F Waste Preparation and Composition

SRS tank farm personnel provided salt cake samples of radioactive waste from Tanks 38H and 46F. The samples originally served for studies of evaporator chemistry. The Savannah River Technology Center (SRTC) received, stored, and processed the samples in a remotely operated, shielded facility. Table II summarizes source information for each sample.

The Tank 38H samples from the evaporator studies included both solids and solutions. The solutions derived from dissolution of solids in water. Aqueous solutions from the second sampling event (July 2001) were combined and evaporated by circulating ambient temperature argon over the surface. After solids crystallized, the solids and solution from the first sampling event (January 2001) were added, additional water was added, the slurry was stirred to saturate the solution and then filtered (0.45 micron nominal pore size). The resulting solution contained 7.1 M sodium ion (density 1.309 g/mL). Dilution with 1.6 M NaOH solution achieved a final salt concentration of 5.6 M sodium ion (density 1.244 g/mL). Table III lists the composition of this solution.

Table II. Origin of Tank 38H and 46F Salt Cake Samples

Source	Tank 38H	Tank 46F
Date Sampled	6/26/01 and 7/25/01	7/3/01
Tank Farm Identification	HTF-395, HTF-396	(no identification)
Approximate volume (L)	34 g solids and 90 mL of solution	~40 g of solids dissolved in 240 mL of water

The Tank 46F samples were all solutions derived from dissolution of solids in water. The solution (240 mL) was evaporated to 100 mL and filtered. The resulting solution contained 4.7 M sodium ion (density 1.167 g/mL). Since the solution was below the desired concentration of 5.6 M Na⁺, it was diluted with Tank 37H/44F composite waste supernate.² Table III lists the composition of the final Tank 46F sample.

Batch Distribution Test Protocol

The protocol for the Cs-137 batch distribution tests included a single extraction step, two scrub steps, and three strip steps (ESS protocol). This differs from previous testing in the number of scrub steps. Previous test results suggested two scrub steps ensured adequate pH adjustment prior to the strip steps.²⁻³ Tests occurred with duplicate samples of each waste solution with targeted organic-to-aqueous (O/A) volume ratios of 0.33 for the extraction step and 5.0 for the scrub and strip steps. We used polypropylene centrifuge tubes (40 mL and 15 mL, Falcon or Corning brand, with lids provided by the manufacturer) for the batch contacts (Figure 1). Aqueous and organic phase volumes were measured gravimetrically based on the solution densities (waste densities

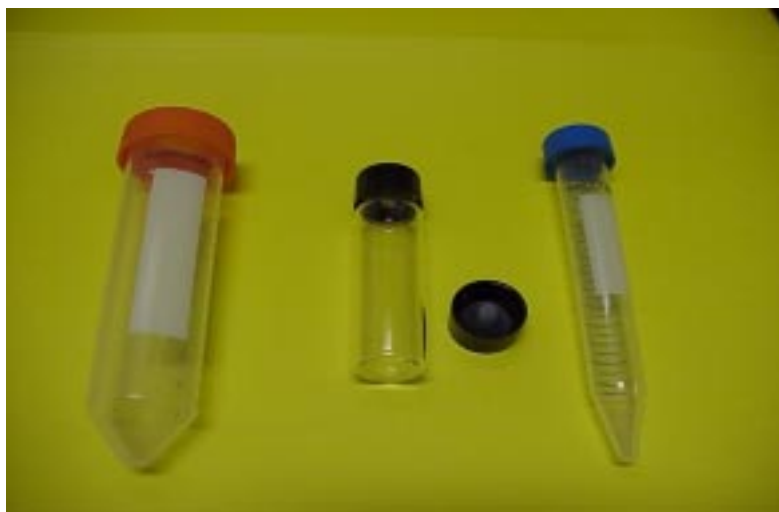
Figure 1. Containers used in batch distribution test protocol

Table III. Composition of the Tank 38H and Tank 46F Samples

<u>Component</u>	<u>Concentration (molar)</u>	
	<u>Tank 38H</u>	<u>Tank 46F</u>
Na ⁺	5.6	5.9
K ⁺	0.0015	0.016
Rb ⁺	<0.002	<0.002
Cs ⁺	2.2x10 ⁻⁶	7.5x10 ⁻⁵
OH ⁻	0.59	3.2
NO ₃ ⁻	0.12	1.28
NO ₂ ⁻	0.060	0.21
AlO ₂ ⁻	0.011	0.20
CO ₃ ²⁻	1.59	0.25
SO ₄ ²⁻	0.35	0.042
Cl ⁻	0.083	0.047
F ⁻	<0.1	<0.1
PO ₄ ³⁻	0.002	0.004
Oxalate	<0.1	<0.1
Formate	<0.2	<0.2
<u>Concentration (mg/L)</u>		
As	<0.5	0.7
Ba	<3	<3
Ca	<1	<1
Cd	<2	<2
Cr	12	26
Hg	4.6	3.0
Mg	<1	<1
Mo	<3	29
Pb	<50	<50
Se	0.9	1.5
Si	87	105
Sr	<1	<1
M-99*	0.23	3.4
M-235	0.25	<.01
M-236	0.06	<.01
M-237	0.16	<.01
M-238	15.	2.6
M-239	<.01	<.01
<u>Concentration (d/m/mL)</u>		
Cs-137	1.97E7	6.99E8
Sr-90	7.4E5	2.9E5
Tc-99	<1.0E4	1.5E5
Pu-238	2 E2	4.8E3
Pu-239/240	7.9E3	5.4E5
Density (g/mL)	1.243	1.238
pH	>13	>14

*Results reflect total mass for the indicated mass number. For example, M-238 includes U-238 and Pu-238.

listed in Table III, solvent density 0.826, and scrub or strip density 1.00). Technicians mixed the phases by shaking 2.5 minutes at ambient cell temperature (25 ± 5 °C) with remote manipulator arms. Centrifuging (~1500 rpm) for 45 minutes separated the phases in the tubes. Using polyethylene eyedroppers, the majority of the organic phase was removed to a clean centrifuge tube. The remaining organic phase and a portion of the aqueous phase were removed and discarded to provide a clean surface for sampling the aqueous phase. Both phases were sampled by 0.5 mL pipette and placed in glass vials. Portions (~0.2 mL) of each phase were gravimetrically diluted in new vials. Aqueous samples were diluted (50-fold for Tank 38H, 130-fold for Tank 46F) using 0.2 M nitric acid. Solvent samples were diluted (same dilution as aqueous samples) using 0.5 M modifier dissolved in Isopar[®] L. Diluted samples from Tank 38H were analyzed by routine gamma spectroscopy methods. To avoid a second dilution of the Tank 46F samples (due to the higher Cs-137 activity compared to Tank 38H), the Tank 46F samples were analyzed in a non-routine geometry that placed the sample vials farther from the detector. The non-routine geometry was calibrated using radioactivity standards. For subsequent steps in the protocol, the remaining solvent was weighed and the amount of aqueous phase calculated.

One of the two initial Tank 38H samples spilled after the first scrub step. At that time, a third sample was extracted and scrubbed. The third sample and the remaining initial sample were processed together through the second scrub and three strip steps.

A pair of centrifuge tubes containing water were handled and centrifuged simultaneously with the waste samples. Temperatures of the water tubes were recorded periodically during the protocol and were used to indicate the actual sample temperatures. It was noted that centrifuging raised the sample temperatures 3 to 5 °C above the cell ambient temperature. Temperatures at the time of the eyedropper phase separation were recorded as the temperature for that distribution coefficient.

Analytical Methods

The SRTC Analytical Development Section provided all analyses by the following routine methods. Al, B, Ba, Ca, Cd, Co, Cr, Cu, Fe, La, Li, Mg, Mn, Mo, Na, Ni, P, Pb, Si, Sn, Sr, Ti, V, Zn, and Zr were measured by ICP-ES. Actinide and fission product isotopes were measured by ICP-MS. Common anions (nitrate, nitrite, sulfate, chloride, fluoride, phosphate, oxalate and formate) were measured by ion chromatography. K, Rb, Cs, Se, As, and Hg were measured by atomic adsorption or emission methods. Cs-137 was measured by gamma spectroscopy using a germanium detector. Sr-90 and Tc-99 were separated and quantified by scintillation counting. Pu isotopes were separated and counted on alpha plates.

Scrub and Strip Solutions

Scrub (0.050 M nitric acid) and strip (0.0010 M nitric acid) solutions were prepared by gravimetric dilution of standardized 0.100 M nitric acid (Fisher, Lot #003535-24) and verified correct by ADS standard titration before use.

Two small bottles of the stock strip solution were placed into the shielded facility, one for use with the Tank 38H sample and one for use with the Tank 46F sample. Following the distribution coefficient tests, the three bottles of strip solution (i.e., the stock solution bottle and the two small bottles that were placed in the cells) were titrated a second time. The pH as indicated by pH paper indicated all samples were approximately pH = 3, as expected for 1 mM nitric acid. ADS titration revealed the strip solution in the bottle used for the Tank 38H test was less than 1 mM. Since this concentration is near the detection limit and the amount of solution was limited, a more accurate estimate could not be made. The initial pH, as indicated by the titration curve, was approximately equal for all three solutions (suggesting that the Tank 38H strip solution was not much lower than 1 mM). Gamma spectroscopy of the contents of the Tank 38H strip solution bottle showed 130 d/m/mL of Cs-137. This level of contamination is not unexpected for a non-radioactive solution that has been used in the Shielded Cells. However, it is less than expected for contamination by a drop of high activity waste (One drop, or 0.05 mL, of the Tank 38H waste would contaminate the solution approximately 100X more than observed). This result suggests the solution was not contaminated by the waste used in the test. Additional analyses of the Tank 38H strip solution showed 1 mM nitrate (as expected), 0.6 mM formate, 0.6 mM chloride, and 0.3 mM sulfate, and 12 mM sodium. The concentrations of the anions identified account for only 33% of the sodium concentration. Because of the pH and titration observations, the remainder of the anion balance cannot be due to hydroxide or carbonate.

Calculations

Distribution coefficients for cesium (D_{Cs}) were calculated by the following equation.

$$D_{Cs} = \frac{[\text{Cs-137 activity in the organic phase}]}{[\text{Cs-137 activity in the aqueous phase}]}$$

The ambient-temperature distribution coefficients were corrected to 25 °C using the following equation and enthalpy (H) values from Bonnesen, et al.¹

$$D_1 = D_0 \exp[(H/R)\{1/T_1 - (1/T_0)\}]$$

$$\text{where } R = 0.0083144 \text{ kJ/mol} \cdot ^\circ\text{K}$$

The enthalpy values used were: extraction, 42.8 kJ/mole; scrub, 61.8 kJ/mole; and strip, 62.5 kJ/mole.

For each set of distribution data, a material balance for Cs-137 was calculated. That is, the activity in the initial solution was compared to sum of the activities in the two final solutions. Distribution coefficients were not calculated for test data in which the material balance did not agree within 15% of the expected total activity. Appendix A lists the measured and calculated results of the experiments.

RESULTS AND DISCUSSION

Dissolved Salt Waste Composition

The Tank 38H sample reflects the potential variability of waste compositions that may arise from salt cake dissolution. The Tank 46F sample shows some of the same variability but to a lesser extent because it was diluted with supernate from Tanks 37F and 44H. The following summarize significant differences between the dissolved salt cake samples and average waste.⁵

- Cesium and potassium concentrations are lower in salt cake.
- Hydroxide and nitrate concentrations are lower in salt cake.
- Carbonate and sulfate concentrations are higher in salt cake.

The Tank 38H sample contained an unusually high amount of carbonate ion. This may result from the sampling method that obtained salt cake from near the surface of the salt deposit. If this layer was ever in contact with air, carbonated salts could have accumulated at the surface from the reaction of carbon dioxide in the air with sodium hydroxide in the waste.

Cs-137 Distribution Coefficients

Table IV summarizes the temperature-corrected distribution coefficients for the two waste samples. The table also includes predicted D_{Cs} values for the extraction step and the D_{Cs} values that form the basis for the CSSX baseline flowsheet.⁵ The predicted values derive from an ORNL model based on data from simulated waste tests. The model predicts values at 25 °C. Strip distribution coefficients for the Tank 38H sample are not reported because the Cs-137 material balance was in error by 60%.

Extraction

The extraction distribution coefficients listed in Table IV exceed the values used in the CSSX flowsheet⁵ and agree reasonably well with predicted values (within 30%). The results indicate that the CSSX process will successfully extract cesium from waste solutions derived from dissolved salt cake. The agreement between real waste and the model in these data is closer than previous comparisons in which they differed by as much as 300%.³ The improvement is likely due to better technique from increased experience in the execution of the ESS protocol in the shielded facility. The remaining discrepancy may be due to the highly unusual composition of the Tank 38H sample.

TABLE IV. Cs-137 Distribution Coefficients for Tank 38H and Tank 46F Waste

Step	Flowsheet Basis	<u>D_{Cs} Values</u>			
		Tank 38H		Tank 46F	
		<u>Predicted</u>	<u>Measured</u>	<u>Predicted</u>	<u>Measured</u>
Extraction	>8	9.6	11.9 ±1.1	22.8	16.6 ±0.5
Scrub 1	>0.6		1.48 ±0.22		1.48 ±0.36
Scrub 2			1.30 ±0.05		1.63 ±0.11
Strip 1	<0.16		--*		0.107 ±0.042
Strip 2			--*		0.071 ±0.011
Strip 3			--*		0.055 ±0.013

*Values not calculated due to poor Cs-137 material balance.

Usually, high-level waste supernates contain mostly monoanions, such as nitrate, nitrite, chloride, and hydroxide. The Tank 38H waste is predominantly dianions (carbonate and sulfate). The model is based on data from simulants containing only mono-anions (hydroxide, nitrate, nitrite, and chloride).

Scrub

The scrub distribution coefficients for both waste samples (1.30 to 1.63) met the requirement of the CSSX flowsheet (>0.6) and are similar to previously measured values for simulants and actual waste (normally 1.0 to 2.0).³⁻⁴

Strip

The strip distribution coefficients for the Tank 46F sample (0.055 to 0.107) compare favorably to the flowsheet basis value of 0.16. The slight downward trend for subsequent strips agrees with previous work with real waste and simulants.³⁻⁴ High values for the strip D_{Cs} values due to alkaline carry-over after a single scrub step reported previously were likely avoided by adding the second scrub step to the ESS protocol.

The strip distribution coefficients for the Tank 38H sample were not calculated because the cesium material balance between the initial and final solutions did not agree. An excess of cesium in the final solutions (aqueous and solvent) compared to the initial solvent solution is likely due to contamination of samples during handling in the Shielded Cells. However, that was not the cause of the imbalance with the Tank 38H sample. Approximately 60% of the starting cesium was missing in the final solutions (Appendix A). The consistency of the imbalance through the three strip steps suggests the cause was not random errors in sampling or dilution. Repetition of the analysis of the organic phase samples proved reproducible to within 15%. However, repeat sampling of the aqueous phases produced highly variable results (±100%). The missing cesium may have sorbed on the sides of the containers or may have been removed when the intermediate (mixed organic and aqueous) phase was discarded. A repetition of the test with particular care in

the sampling and analysis of the intermediate phase and solution containers could evaluate these possibilities.

Problems with material balance can indicate formation of a small amount of a third phase enriched in cesium. However, the known initiators of third-phase formation (high cesium or potassium and low temperature) were not present in this test.⁴ Although it is possible that the problem derived from the unusual composition of the Tank 38H sample, this seems unlikely since there is no obvious mechanism. That is, the problem in the strip step is separated from the Tank 38H sample by two scrub steps.

The strip solution used with the Tank 38H sample was subsequently found to have been contaminated during use, yielding high sodium and low acid concentrations compared to normal strip solution. Although a strip solution containing high sodium ion concentration or the wrong acid concentration would adversely affect the strip results, it is unlikely that it would affect the material balance. The contaminants identified in the contaminated solution do not suggest a mechanism for the poor material balance. However, since the anion balance is incomplete, it is possible that the unidentified component may have affected the cesium balance.

CONCLUSIONS

Researchers characterized high-level waste solutions derived from dissolved salt cake from the SRS tank farm and measured the cesium distribution coefficients for extraction, scrub, and stripping steps of the CSSX flowsheet. The measurements used two SRS high-level waste samples derived entirely or in part from salt cake. A sample of Tank 38H salt cake provided a waste solution suitable for testing (5.6 M Na⁺). A dilute solution of Tank 46F salt cake was combined with concentrated supernate solution from Tank 37H/44F to achieve the desired salt concentration for testing. The chemical compositions of both samples differed significantly from average waste, but the differences were largely predictable from solubility considerations. Extraction and scrub D_{Cs} values for both samples exceeded process requirements and agreed well with model predictions. Strip D_{Cs} values for the Tank 46F sample also met process requirements. However, strip D_{Cs} values could not be calculated for the Tank 38H sample due to the poor material balance for Cs-137. Contamination during sample handling within the shielded facility did not cause the cesium material balance problem. Sorption on container walls, removal in a waste cut, or the unusual waste composition may explain the discrepancy.

FUTURE WORK

The following additional investigations are recommended to resolve question of the poor Tank 38H strip material balance and to improve understanding of the CSSX process chemistry.

- Repeat the ESS protocol with Tank 38H salt cake waste. The protocol should be repeated with emphasis on obtaining a good Cs-137 material balance during the strip steps. In particular, the intermediate layer normally discarded to waste should be retained and analyzed if the material balance is in error.
- The ORNL model for extraction D values should be improved to accommodate a broader range of anions, particularly dianions. This should provide better agreement between predicted and measured extraction D_{Cs} values.

ACKNOWLEDGMENTS

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APPENDIX A **Measured and Calculated Results**

Tank 38 H Sample							
Step			Sample 1	Sample 2	Sample 3	Average	Error
Extraction							
Experimental data	Temp (deg C)		23.6	23.6	25.9		
	Aqueous activity (d/m/mL)		3.68E+06	3.60E+06	3.58E+06		
	Organic activity (d/m/mL)		4.42E+07	4.47E+07	4.46E+07		
	D-value		12.0	12.4	12.5		
D-value corrected to 25 deg C			11.1	11.4	13.1	11.9	1.1
Cs-137 material balance	Initial (d/m)		5.91E+08	5.91E+08	5.91E+08		
	Final (d/m)		5.52E+08	5.55E+08	5.53E+08		
Scrub #1							
Experimental data	Temp (deg C)		27.6	25.2	26.5		
	Aqueous activity (d/m/mL)		2.77E+07	2.67E+07	3.48E+07		
	Organic activity (d/m/mL)		3.78E+07	3.91E+07	3.83E+07		
	D-value		1.36	1.46	1.10		
D-value corrected to 25 deg C			1.69	1.49	1.25	1.48	0.22
Cs-137 material balance	Initial (d/m)		4.42E+08	4.47E+08	4.46E+08		
	Final (d/m)		4.33E+08	4.44E+08	4.53E+08		
Scrub #2							
Experimental data	Temp (deg C)		21.3		21.3		
	Aqueous activity (d/m/mL)		1.89E+07		1.97E+07		
	Organic activity (d/m/mL)		3.46E+07		3.40E+07		
	D-value		1.83		1.73		
D-value corrected to 25 deg C			1.34		1.26	1.30	0.05
Cs-137 material balance	Initial (d/m)		3.78E+08		3.83E+08		
	Final (d/m)		3.84E+08		3.79E+08		
Strip #1							
Experimental data	Temp (deg C)		24.3		23.1		
	Aqueous activity (d/m/mL)		6.43E+07		1.89E+07		
	Organic activity (d/m/mL)		1.10E+07		1.05E+07		
	D-value						
D-value corrected to 25 deg C					Error in material balance.		
Cs-137 material balance	Initial (d/m)		3.46E+08		3.40E+08		
	Final (d/m)		2.39E+08		1.43E+08		
Strip #2							
Experimental data	Temp (deg C)		24.8		24.2		
	Aqueous activity (d/m/mL)		4.46E+06		5.49E+06		
	Organic activity (d/m/mL)		3.07E+06		2.81E+06		
	D-value						
D-value corrected to 25 deg C					Error in material balance.		
Cs-137 material balance	Initial (d/m)		1.10E+08		1.05E+08		
	Final (d/m)		3.96E+07		3.91E+07		
Strip #3							
Experimental data	Temp (deg C)		26.5		25.3		
	Aqueous activity (d/m/mL)		1.31E+06		1.24E+06		
	Organic activity (d/m/mL)		7.65E+05		6.46E+05		
	D-value						
D-value corrected to 25 deg C					Error in material balance.		
Cs-137 material balance	Initial (d/m)		3.07E+07		2.81E+07		
	Final (d/m)		1.03E+07		8.94E+06		

Tank 46F Sample						
Step		Sample 1	Sample 2	Average	Error	
Extraction						
Experimental data	Temp (deg C)	22.6	22.6			
	Aqueous activity (d/m/mL)	9.20E+07	9.54E+07			
	Organic activity (d/m/mL)	1.79E+09	1.78E+09			
	D-value	19.5	18.7			
D-value corrected to 25 deg C		16.9	16.2	16.6	0.5	
Cs-137 material balance	Initial (d/m)	2.10E+10	2.10E+10			
	Final (d/m)	2.07E+10	2.07E+10			
Scrub #1						
Experimental data	Temp (deg C)	25.8	23.4			
	Aqueous activity (d/m/mL)	1.03E+09	1.06E+09			
	Organic activity (d/m/mL)	1.67E+09	1.48E+09			
	D-value	1.62	1.40			
D-value corrected to 25 deg C		1.73	1.22	1.48	0.36	
Cs-137 material balance	Initial (d/m)	1.79E+10	1.78E+10			
	Final (d/m)	1.88E+10	1.69E+10			
Scrub #2						
Experimental data	Temp (deg C)	27.2	25.7			
	Aqueous activity (d/m/mL)	1.02E+09	8.20E+08			
	Organic activity (d/m/mL)	1.32E+09	1.32E+09			
	D-value	1.29	1.61			
D-value corrected to 25 deg C		1.55	1.71	1.63	0.11	
Cs-137 material balance	Initial (d/m)	1.67E+10	1.48E+10			
	Final (d/m)	1.52E+10	1.48E+10			
Strip #1						
Experimental data	Temp (deg C)	29.1	23.3			
	Aqueous activity (d/m/mL)	4.01E+09	4.62E+09			
	Organic activity (d/m/mL)	3.88E+08	4.13E+08			
	D-value	0.097	0.089			
D-value corrected to 25 deg C		0.136	0.077	0.107	0.042	
Cs-137 material balance	Initial (d/m)	1.32E+10	1.32E+10			
	Final (d/m)	1.19E+10	1.34E+10			
Strip #2						
Experimental data	Temp (deg C)	27.4	24.7			
	Aqueous activity (d/m/mL)	1.62E+09	1.51E+09			
	Organic activity (d/m/mL)	1.04E+08	9.75E+07			
	D-value	0.064	0.065			
D-value corrected to 25 deg C		0.079	0.063	0.071	0.011	
Cs-137 material balance	Initial (d/m)	3.88E+09	4.13E+09			
	Final (d/m)	4.28E+09	4.00E+09			
Strip #3						
Experimental data	Temp (deg C)	27.5	25.3			
	Aqueous activity (d/m/mL)	4.02E+08	3.87E+08			
	Organic activity (d/m/mL)	2.09E+07	1.75E+07			
	D-value	0.052	0.045			
D-value corrected to 25 deg C		0.064	0.046	0.055	0.013	
Cs-137 material balance	Initial (d/m)	1.04E+09	9.75E+08			
	Final (d/m)	1.01E+09	9.49E+08			