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# **CHEMICAL OXYGEN DEMAND (COD) FOR MONITORING REDUCTION-OXIDATION (REDOX) EQUILIBRIUM DURING HIGH LEVEL WASTE (HLW) VITRIFICATION (U)**

**C. M. Jantzen and M.J. Whitaker**

**Publication Date: April 30, 2004**

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## EXECUTIVE SUMMARY

High-level nuclear waste is being immobilized at the Savannah River Site (SRS) by vitrification into borosilicate glass at the Defense Waste Processing Facility (DWPF). Control of the REDuction/OXidation (REDOX) equilibrium in the DWPF melter is critical for processing high level liquid wastes. Based upon previous research, an acceptable iron REDOX ratio was defined for the DWPF melts as  $0.09 \leq \text{Fe}^{2+}/\Sigma\text{Fe} \leq 0.33$ . Controlling the DWPF melter at a REDuction/OXidation (REDOX) equilibrium of  $\text{Fe}^{+2}/\Sigma\text{Fe} \leq 0.33$  prevents the potential for metallic and metallic sulfide species to form and accumulate on the floor of the melter. Control of foaming due to deoxygenation of manganic species is achieved by converting 66-100% of the  $\text{MnO}_2$  or  $\text{Mn}_2\text{O}_3$  species in a waste feed to  $\text{MnO}$  before the waste is fed to the DWPF melter. At the lower redox limit of  $\text{Fe}^{+2}/\Sigma\text{Fe} \sim 0.09$  about 99% of the  $\text{Mn}^{+4}/\text{Mn}^{+3}$  is converted to  $\text{Mn}^{+2}$ . Therefore, the lower REDOX limit eliminates melter foaming from deoxygenation.

Organic and nitrate concentrations in the DWPF melter feed are the major parameters influencing melt REDOX. Organics such as formates act as reductants while nitrates, nitrites, and manganic ( $\text{Mn}^{+4}$  and  $\text{Mn}^{+3}$ ) species act as oxidants. During melting, the REDOX of the melt pool cannot be measured. Therefore, the  $\text{Fe}^{+2}/\Sigma\text{Fe}$  ratio in the glass poured from the melter must be related to melter feed organic and oxidant concentrations to ensure production of a high quality glass without impacting production rate (e.g., from foaming) or melter life (e.g., from metal formation and accumulation).

A new Electron Equivalents REDOX model has been developed for DWPF so that the  $\text{Fe}^{+2}/\Sigma\text{Fe}$  ratio can be calculated from feed reductants (formic acid, coal, soluble oxalate, and insoluble oxalate), feed oxidants (nitrate, nitrite, soluble  $\text{Mn}^{+4}$ , and insoluble  $\text{Mn}^{+4}$ ), and wt% solids. This requires nine different analyses to be performed. Four of the analyses involve carbon in a variety of oxidation states. The most difficult analysis of a carbon containing species is that of coal. Because of the difficulty and time consuming nature of this analysis it may only be performed once per sludge batch and assumed representative of an entire sludge batch.

A simple analytic technique, known as Chemical Oxygen Demand (COD), was investigated to replace the four carbon species analyses which would leave five analyses to be performed (nitrate, nitrite, soluble manganese, insoluble manganese, and wt% solids) in order to calculate the  $\text{Fe}^{+2}/\Sigma\text{Fe}$  ratio using the Electron Equivalents REDOX model for DWPF. During the course of this study it was determined that the COD response actually measured the sum of all the electron transfers, e.g. the relative contributions of each of the reductants and oxidants. Therefore, only two analyses appear to be needed to predict DWPF REDOX, e.g. COD and SRAT wt% solids. Fewer analyses will improve the feed acceptability turn around time by substituting a simpler analysis, COD, for multiple analyses, some of which are difficult to perform. Substitution of the COD technique will not compromise the REDOX control strategy currently in use in DWPF. Use of the COD technique is, therefore, considered promising but needs further refinement before it can be implemented in DWPF.

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**LIST OF ACRONYMS**

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AVE	Average
COD	Chemical Oxygen Demand
DWPF	Defense Waste Processing Facility
EE	Electron Equivalents
HLW	High Level Waste
LSFM	Large Slurry Fed Melter
PNNL	Pacific Northwest National Laboratory
PSCM	Pilot Scale Ceramic Melter
REDOX	REDuction/OXidation
RSD	Relative Standard Deviation
SD	Standard Deviation
SME	Slurry Mix Evaporator
SRAT	Sludge Receipt and Adjustment Tank
SRL	Savannah River Laboratory
SRS	Savannah River Site
SRTC	Savannah River Technology Center
WSRC	Westinghouse Savannah River Co.

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REDUCTION-OXIDATION (REDOX) EQUILIBRIUM DURING  
HIGH LEVEL WASTE (HLW) VITRIFICATION (U)**

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**1.0 INTRODUCTION**

**1.1 DWPF PROCESS CONSTRAINTS**

High-level nuclear waste is being immobilized at the Savannah River Site (SRS) by vitrification into borosilicate glass at the Defense Waste Processing Facility (DWPF). In the DWPF Sludge Receipt and Adjustment Tank (SRAT), the insoluble fraction of the waste sludge is refluxed with acid [1] for the following reasons:

- control potential foaming by gaseous species in the melter by:
  - converting  $\text{NO}_3$  (entering as nitrate species in the feed) to  $\text{NO}_2$ ,  $\text{NO}$ , and/or  $\text{N}_2$ ,
  - converting carbonates in the feed to  $\text{CO}_2$ , and
  - converting >66% of the oxidized  $\text{Mn}^{+4}$  or  $\text{Mn}^{+3}$  present as  $\text{MnO}_2$ ,  $\text{Mn}_2\text{O}_3$ , and  $\text{Mn}_3\text{O}_4$  and/or hydrous complexes in the feed to  $\text{MnO}$ , liberating  $\text{O}_2$ ;
- steam strip mercury for subsequent removal,  $\text{HgO} \rightarrow \text{Hg}^0$ ; and
- improve slurry rheology by neutralizing excess hydroxide ( $\text{OH}$ ) in the feed.

The SRAT product is then fed to the DWPF Slurry Mix Evaporator (SME), where a borosilicate glass frit slurry is added to produce the melter feed slurry. The melter feed slurry is typically concentrated to 45-50 wt% total solids in the SME and then fed to the DWPF joule-heated melter where it is fused into glass (vitrified) at 1150°C.

Prior to 1982, the DWPF reference flowsheet prescribed that 28 wt% sludge oxides (on a dry calcine basis) from the SRAT be combined with 72 wt% frit oxides in the SME. Formic acid was the only acid added during SRAT processing. Most glasses produced from this "sludge-only" and "formic acid-only" flowsheet were highly reduced, e.g.  $\gg 0.33 \text{ Fe}^{+2}/\Sigma\text{Fe}$ . Any carbon containing species in melter feeds (coal, oxalate, formic acid, sugar) cause reduction of transition metal species such as  $\text{Fe}^{+3}$  and  $\text{Mn}^{+4}$  at the elevated melter temperatures [2,3,4]. The interaction of the carbon with the transition metal species in the waste feed occurs primarily in the melter cold cap [5,6].

Excess reduction of transition metal species in the melter can cause the following to occur:

- liberation of oxygen which can cause foaming from decomposition of  $Mn^{+4}$  or  $Mn^{+3}$  species if they were not previously reduced during SRAT processing
- reduction of metallic species such as  $NiO \rightarrow Ni^{\circ}$  and  $RuO_2 \rightarrow Ru^{\circ}$  which may fall to the melter floor and cause shorting of electrical pathways in the melt and accumulations which may hinder glass pouring
- reduction of sulfate ( $SO_4^{\bar{}}$ ) to sulfide ( $S^{\bar{}}$ ) which can complex with  $Ni^{\circ}$  and/or  $Fe^{\circ}$  to form metal sulfides which can fall to the melter floor and cause shorting of electrical pathways and/or hinder glass pouring
- reduced glasses which can be less durable than their oxidized equivalents [7].

Controlling the DWPF melter at a REDuction/OXidation (REDOX) equilibrium of  $Fe^{+2}/\Sigma Fe \leq 0.33$  [2, 8] prevents the potential for conversion of  $NiO \rightarrow Ni^{\circ}$ ,  $RuO_2 \rightarrow Ru^{\circ}$ , and  $2SO_4^{\bar{}} \rightarrow S_2 + 4O_2$  during vitrification. Control of foaming due to deoxygenation of manganic species is achieved by having 66-100% of the  $MnO_2$  or  $Mn_2O_3$  species converted to  $MnO$  [9] during SRAT refluxing. At the lower redox limit of  $Fe^{+2}/\Sigma Fe \sim 0.09$  about 99% of the  $Mn^{+3}$  is converted to  $Mn^{+2}$  [2, 8]. Therefore, the lower REDOX limit eliminates melter foaming from deoxygenation.

While nitric acid can be used to control feed rheology and destroy carbonates, only a reducing acid such as formic acid can convert  $HgO \rightarrow Hg^0$  when it is present in a feed and convert  $MnO_2 \rightarrow MnO + \frac{1}{2} O_2$  in the SRAT. The REDOX equilibrium in DWPF is controlled by balancing organic reductants (formic acid, coal, soluble oxalates and insoluble oxalates) and oxidizers (nitric acid, sodium nitrate, sodium nitrite, and  $MnO_2$ ) additions to the SRAT.

Chemical reduction is defined as making an atom or molecule less positive by electron transfer, while oxidation is defined as making an atom or molecule more positive by electron transfer. Therefore, the number of electrons transferred for each REDuction/OXidation reaction can be summed and an Electron Equivalents term for each organic and oxidant species defined.‡ The formalism developed [10] for DWPF REDOX control from feed composition is represented by the number of electrons gained during reduction of an oxidant or lost during oxidation of a reductant. The overall relationship between the REDOX ratio and the feed is given in term of the transfer of molar Electron Equivalents,  $\xi$ :

---

‡ Because the water content of a melter feed alters the species concentrations of the [reductants] and [oxidants], it can influence the equilibrium oxygen fugacity ( $f_{O_2}$ ) in a melter during vitrification.

Since the effects of water on oxygen fugacity are small relative to the impact of dilution on feed concentrations, the molar concentrations of the reductants and oxidants must be adjusted to a constant wt% solids basis. A value of 45 wt% solids was chosen as the nominal wt% solids for most DWPF melter feeds in the Electron Equivalents model shown in Equation 1.

**Equation 1** 
$$\frac{Fe^{2+}}{\Sigma Fe} = f \left[ \left( 2[F] + 4[C] + 4[O_T] - 5[N] - 2[Mn] \right) \frac{45}{T} \right] = f[\xi]$$

where  $f$  = indicates a function  
 $[F]$  = formate (mol/kg feed)  
 $[C]$  = coal (carbon) (mol/kg feed)  
 $[O_T]$  = oxalate<sub>Total</sub> (soluble and insoluble) (mol/kg feed)  
 $[N]$  = nitrate + nitrite (mol/kg feed)  
 $[Mn]$  = manganese<sub>Total</sub> (soluble and insoluble) (mol/kg feed)  
 $T$  = total solids (wt%)

and  $\xi$  (mol/kg feed at 45 wt% solids) =  $(2[F] + 4[C] + 4[O_T] - 5[N] - 2[Mn]) \frac{45}{T}$

## 1.2 PURPOSE OF THE CURRENT STUDY

In order to implement the Electron Equivalents REDOX model in DWPF, nine\* analyses are needed on the DWPF feed (formate, nitrate, nitrite, soluble oxalate, insoluble oxalate, coal, soluble Mn<sup>+2</sup>, insoluble Mn<sup>+4</sup>, and wt% solids). The current study addresses a simple and rapid analysis method that would substitute one measurement, chemical oxygen demand (COD) for four of the analyses (formate, soluble oxalate, insoluble oxalate, and coal). In addition, COD analyses may monitor the relative role of both the reductants and the oxidants which would result in even fewer complex analyses to perform. This would significantly improve melter feed acceptability turn around time. In addition, the method for coal is difficult and time consuming, involving sieving and isolation of the coal from the remainder of the sludge by wet sieving [11]. Substitution of a simpler method to measure the sum of the organic reductants and/or the oxidants would improve the overall feed acceptability turn around time without compromising REDOX control in DWPF.

## 2.0 BACKGROUND

### 2.1 REDOX PROCESS LIMITS

The DWPF melt and glass REDOX equilibria have been well studied. Based on the work of Schreiber[2] and Goldman[3], Jantzen and Plodinec [8] originally defined acceptable REDOX ratios (based on the measured  $Fe^{2+}/\Sigma Fe$  ratio) for any DWPF-type melt to be greater than or equal to 0.09 (to prevent foaming via the deoxygenation of MnO<sub>2</sub>, Mn<sub>2</sub>O<sub>3</sub>, and Mn<sub>3</sub>O<sub>4</sub> to MnO) and less than or equal to 0.33 (to prevent metallic nickel and nickel sulfide formation). Formate and nitrate concentrations in the melter feed appear to be the

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\* seven analyses if total oxalate (soluble and insoluble) and total manganese (soluble and insoluble) can be measured together

major parameters influencing melt REDOX during vitrification. The formate is a reductant while the nitrate is an oxidant. Since the melt REDOX ratio cannot be directly measured during processing,<sup>†</sup> the melt REDOX ratio has been related to feed reductant (formate) and oxidizer (nitrate) concentrations (which can be measured). Thus the proper balance between reductants and oxidizers ensures production of a high quality glass, free of metal sulfide precipitation, without impacting production rate (e.g., from foaming), melter life (e.g., from metal formation), or glass performance (e.g. from degradation of durability).

## 2.2 REDOX MODELING

Initially a REDOX model with 1:1 stoichiometry of a single reductant (formic acid) and a single oxidant (nitric acid) was developed known as the {[F]-[N]} model [12, 13, 14, 15]. In 1997, the data used to develop the {[F]-[N]} relationship was revisited and glass quality and REDOX measurement criteria were developed to screen the experimental data used for modeling [16]. This redefined the population of glasses being modeled by excluding those below the  $\text{Fe}^{+2}/\Sigma\text{Fe}$  measurement detection limit of 0.03 and those that precipitated metallic and/or sulfide species. Regression of the redefined data demonstrated that the {[F]-[N]} parameter was a less accurate predictor ( $R^2=0.68$ ) of waste glass REDOX than had previously been calculated ( $R^2=0.80$ ). In addition, there was no known mechanistic impetus for using the molar difference of the reductants and oxidants for REDOX prediction. This artificially set the relative oxidation/reduction potentials of nitrate and formate to be equivalent when it was well known that nitric acid is a strong oxidizer and formic acid is a weak reductant.

The regression of the redefined data showed that there was an {[F]-3[N]} relationship between the feed reductants and oxidants and the REDOX ratio of the glass yielding an  $R^2=0.88$ . The {[F]-3[N]} model, based solely on the molar formate and nitrate concentrations in the DWPF feeds, was implemented in DWPF in January 2000 while processing Sludge Batch 1B (SB1B, SRAT Batch 134). The {[F]-3[N]} REDOX model has been used since then to control REDOX at a target  $\text{Fe}^{2+}/\Sigma\text{Fe} \sim 0.2$  in an effort to improve melt rate by reducing the foaming caused by the excess nitric acid being added.

Both the {[F]-[N]} and the {[F]-3[N]} REDOX models assumed that the melter feeds were properly formated and refluxed to ensure that 66-100% of the  $\text{Mn}^{3+}$  and  $\text{Mn}^{4+}$  were converted to  $\text{Mn}^{+2}$  as  $\text{Mn}(\text{COOH})_2$ . They were each based on only one reductant (formate) and one oxidant (nitrate) even though other organics such as phenol were present and manganic species were present. The recently developed Electron Equivalents REDOX model [10] makes no assumption about the oxidation state of the manganese or about the total destruction of the nitrite during SRAT processing. In addition, the Electron Equivalents REDOX model accounts for additional reductants such as oxalate and coal (see Equation 1) and the expanded version of this model also accounts for the

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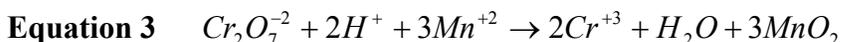
<sup>†</sup> Originally the REDOX for a vitrified melter feed sample was to be measured to determine feed acceptability; however, the expense and time necessary for these measurements provided impetus to relate melt REDOX to feed chemistry.

use of solid reductants such as sugar. The Electron Equivalents REDOX model has an  $R^2 = 0.80$  [10].

### 2.3 CHEMICAL OXYGEN DEMAND (COD) MEASUREMENTS

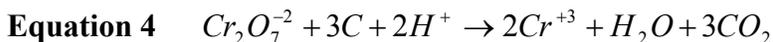
Chemical Oxygen Demand (COD) is a titrimetric technique that analyzes the reducing strength of melter feeds. The COD technique for glass was developed by the FMC Corporation [17]. It is a minor variation of the ASTM Standard Method for Chemical Oxygen Demand of Industrial Waste Water which measures the oxidizable organic matter content in water [18]. The COD method uses a strong chemical oxidant in acid solution and heat to oxidize organic carbon to  $CO_2$  and  $H_2O$ . By definition, chemical oxygen demand is “a measure of the oxygen equivalent of the organic matter content of a sample that is susceptible to oxidation by a strong chemical oxidant [18].”

For application to melter feeds, a known mass of melter feed is refluxed for several hours with sulfuric acid and potassium dichromate. Any oxidizable component in the feed will be oxidized under these conditions. The only reactions which should occur with the melter feeds are those of the dichromate with organics and any reduced species such as  $Mn^{+2}$  in the feed via reactions such as the following:



The rest of the components in the feed, including all of the oxides in the frit, are already in their highest oxidation state. If there is no  $Mn^{+2}$  in the feed, only the organic contribution will be measured by COD analysis.

The dichromate is added in excess, and this excess is titrated with  $Fe^{+2}$ . Since the amount of dichromate initially added is known, as is the amount left after reflux, the amount consumed by the feed is found by difference. The results are expressed as weight percent carbon equivalent, i.e. the amount of carbon that would be consumed by that amount of dichromate:



The COD measurement technique was explored at SRTC (then known as the Savannah River Laboratory, SRL) in 1982 by P.D. Soper [17]. At that time a formic acid only and sludge only flowsheet were being developed for the DWPF. The precision of the COD measurement technique was found to be excellent. In addition, the measured COD response was determined to be a linear function of the REDOX ratio expressed as  $Fe^{+2}/Fe^{+3}$ . Both crucible and pilot scale melter feeds from SRL (Large Slurry Fed Melter, LSFM) and Pacific Northwest Laboratory (PNL Small Cylindrical Melter, PSCM) were tested. It was noted that the COD vs.  $Fe^{+2}/Fe^{+3}$  relationship was melter specific due to

larger air in-leakage of the SRL LSFM versus the PSCM, but that the COD method could be useful to control the reducing capacity of the feed independent of melter air in-leakage.

This method calculates the carbon equivalents in grams per mole of reductant and adjusts them for their relative reducing power. For example in one mole of formate there is only one mole of carbon which has an atomic weight of 12 but a valence of +2. In one mole of coal there is also one mole of carbon, it has an atomic weight of 12, and a valence of 0. Due to the difference in the oxidation states of the carbon in coal and formic acid, a mole of carbon in formic acid transfers 2 electrons when it oxidizes to  $C^{+4}O_2$  while the coal transfers 4 electrons when it oxidizes to  $C^{+4}O_2$ . Therefore, one mole of carbon in formic acid is considered as 6 grams of carbon equivalents per mole on COD analysis while one mole of carbon in coal is considered as 12 grams of carbon equivalents per mole [17]. Moreover, Soper [17] had normalized the carbon equivalents by the wt% feed solids in order to be able to model both dilute and concentrated feeds. Therefore, Soper [17] expressed the COD relationship to  $Fe^{+2}/Fe^{+3}$  in terms of wt%C/feed solids.

Investigation of the COD methodology to measure REDOX of feeds in DWPF was abandoned in 1986 when high nitrate compounds such as sodium nitrate (a corrosion inhibitor)[19], hydroxylamine nitrate (to reduce nitrite to nitrous oxide), and nitric acid (substitution of nitric acid for a portion of the formic acid) [20] were introduced into various flowsheet options. It was assumed that the presence of excessive nitrate, a melt oxidizer, would interfere with the COD measurement.

### 3.0 EXPERIMENTAL

#### 3.1 FEED PREPARATION AND ANALYSES

Detailed preparation and analyses of the SME feeds used in the COD measurements and the crucible studies are given elsewhere [10]. The feeds varied in formic, nitrate, nitrite, oxalate, coal, and manganese content over the concentrations anticipated for SB3. The feeds all contained between 10 wt% to 100 wt% of the amount of noble metals calculated to be present in Tank 7. Manganese in the air-dried simulated sludges was 2.92 wt% on a dry solids basis (110-115°C) for the Tank 8 simulants\* and 3.87 wt% on a dry solids basis (110-115°C) in the Tank 7 simulant.

Coal size, type and treatment were varied as shown in Table 1. Coal treatment included irradiation and caustic treatment of filter specification size coal (coarse) and fine coal to simulate degradation after years of tank storage. The filter specification size coal was

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\* A Tank 8 simulant was used until a SB3 Tank 7 simulant was available. The elemental composition of both was similar and the anion content of the Tank 8 simulant was adjusted to match the anticipated composition of Tank 7.

provided by the manufacturer of the K-Area filters and ranged between 0.6 to 0.8mm. The fine coal was ~ 0.21mm in diameter.<sup>†</sup>

Table 1 summarizes the measured REDOX and waste loading data for the 53 glasses used in the development of the Electron Equivalents model. Table 1 also gives the SME concentrations on which COD was measured.

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<sup>†</sup> J.R. Fowler, "Particle Distribution of Coal and Sand in Tank 7," Interoffice Memorandum, September 20, 1979.

**Table 1 Adjusted SME Concentrations, Chemical Oxygen Demand, REDOX, and Calculated Electron Equivalents**

Sample ID	COD (mol/kg)	COD (mol/kg) * SRAT Wt% Solids	Coal Type and Size	Target Waste Loading (wt%)	REDOX Fe <sup>+2</sup> /ΣFe	Adjusted SME Total Solids (wt%)	Adjusted Oxalate (mol/kg)	Adjusted Formate (mol/kg)	Adjusted Nitrate (mol/kg)	Adjusted Coal (mol/kg)	Adjusted Mn (mol/kg)	Electron Equivalents (mol/kg feed @ 45 wt% solids)
SB3-1-25-320	0.21	5.52	None	25	0.119	42.80	0.000	0.321	0.154	0.000	0.065	-0.271
SB3-1-30-320	0.22	5.32	None	30	0.074	40.47	0.000	0.322	0.154	0.000	0.098	-0.360
SB3-5-30-320	0.27	8.83	None	30	0.407	43.24	0.270	0.385	0.187	0.000	0.064	0.820
SB3-5-35-320	0.35	11.06	None	35	0.451	41.06	0.298	0.401	0.195	0.000	0.074	0.956
SB3-6-25-320	0.23	7.44	None	25	0.370	46.66	0.128	0.360	0.139	0.000	0.061	0.400
SB3-6-30-320	0.29	8.68	None	30	0.267	42.01	0.155	0.375	0.145	0.000	0.076	0.527
SB3-7-25-320	0.23	6.17	None	25	0.125	42.01	0.237	0.304	0.265	0.000	0.049	0.141
SB3-7-30-320	0.29	7.34	None	30	0.193	37.75	0.284	0.315	0.275	0.000	0.063	0.317
SB3-7-35-320	0.31	7.57	None	35	0.142	36.02	0.318	0.325	0.283	0.000	0.073	0.446
SB3-15-30-320	0.38	9.97	Fine, Caustic Treated, Irradiated	30	0.331	39.04	0.197	0.573	0.248	0.006	0.057	0.702
SB3-1-25-202	0.24	6.43	None	25	0.105	42.84	0.000	0.321	0.154	0.000	0.065	-0.271
SB3-1-35-202	0.26	6.03	None	35	0.118	35.50	0.000	0.362	0.174	0.000	0.074	-0.369
SB3-2-25-200	0.23	6.10	Act Carb	25	0.242	42.99	0.000	0.307	0.159	0.078	0.061	0.006
SB3-2-30-200	0.23	5.64	Act Carb	30	0.212	37.27	0.000	0.325	0.169	0.089	0.070	0.027
SB3-2-35-200	0.25	6.02	Act Carb	35	0.136	33.67	0.000	0.345	0.179	0.090	0.071	0.017
SB3-3-25-200	0.27	7.59	Act Carb	25	0.275	43.85	0.000	0.399	0.162	0.078	0.059	0.189
SB3-4-25-202	0.23	6.17	Act Carb	25	0.138	43.24	0.000	0.302	0.164	0.079	0.063	-0.027
SB3-6-25-202	0.28	8.74	None	25	0.336	46.67	0.128	0.359	0.139	0.000	0.061	0.400
SB3-6-30-202	0.28	8.35	None	30	0.288	42.66	0.161	0.371	0.144	0.000	0.080	0.536
SB3-7-25-202	0.28	7.49	None	25	0.193	42.24	0.281	0.291	0.254	0.000	0.065	0.327
SB3-7-30-202	0.34	8.65	None	30	0.141	38.15	0.253	0.324	0.283	0.000	0.052	0.166
SB3-7-35-202	0.27	6.54	None	35	0.155	35.40	0.307	0.328	0.286	0.000	0.070	0.400
SB3-8-25-202	0.28	7.51	Act Carb	25	0.272	41.79	0.234	0.299	0.266	0.059	0.048	0.373
SB3-8-35-202	0.32	7.77	Act Carb	35	0.175	34.91	0.262	0.334	0.297	0.066	0.053	0.500

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Sample ID	COD (mol/kg)	COD (mol/kg) * SRAT Wt% Solids	Coal Type and Size	Target Waste Loading (wt%)	REDOX Fe <sup>+2</sup> /ΣFe	Adjusted SME Total Solids (wt%)	Adjusted Oxalate (mol/kg)	Adjusted Formate (mol/kg)	Adjusted Nitrate (mol/kg)	Adjusted Coal (mol/kg)	Adjusted Mn (mol/kg)	Electron Equivalents (mol/kg feed @ 45 wt% solids)
SB3-9-25-202	0.32	8.90	Spec Size	25	0.243	42.89	0.227	0.327	0.219	0.059	0.051	0.631
SB3-9-30-202	0.40	10.43	Spec Size	30	0.225	39.02	0.243	0.349	0.234	0.063	0.055	0.740
SB3-10-35-202	0.41	10.07	Spec Size	35	0.342	35.71	0.208	0.412	0.227	0.007	0.056	0.549
SB3-11-25-202	0.37	10.39	Spec Size	25	0.387	42.79	0.180	0.409	0.189	0.006	0.051	0.539
SB3-12-25-202	0.27	7.59	Spec Size, Caustic Treated, Irradiated	25	0.334	43.03	0.186	0.424	0.237	0.059	0.051	0.562
SB3-12-30-202	0.41	11.04	Spec Size, Caustic Treated, Irradiated	30	0.346	39.21	0.198	0.452	0.253	0.063	0.055	0.658
SB3-13-25-202	0.30	8.72	Fine	25	0.342	43.15	0.202	0.486	0.254	0.059	0.053	0.670
SB3-13-30-202	0.36	9.57	Fine	30	0.352	39.31	0.215	0.519	0.271	0.063	0.057	0.785
SB3-14-30-202	0.42	11.07	Fine	30	0.371	39.01	0.219	0.440	0.236	0.007	0.056	0.568
SB3-15-25-202	0.27	7.77	Fine, Caustic Treated, Irradiated	25	0.382	42.87	0.185	0.537	0.232	0.006	0.053	0.600
SB3-15-30-202	0.34	9.13	Fine, Caustic Treated, Irradiated	30	0.360	39.04	0.197	0.573	0.248	0.006	0.057	0.703
SB3-16-25-202	0.27	7.42	Fine, Caustic Treated, Irradiated	25	0.321	42.52	0.212	0.412	0.249	0.059	0.052	0.588
SB3-18-25-202	0.21	5.73	Fine, Caustic Treated, Irradiated	25	0.044	43.32	0.000	0.414	0.221	0.078	0.066	-0.104

### 3.2 COD ANALYSES

#### 3.2.1 Standard COD Analyses

Over 200 simulated SB3 SME products, each representative of a glass sample made during the development of the Electron Equivalents REDOX model, were reacted with an acidic solution of potassium dichromate in the presence of a catalyst (silver) and digested at a temperature of 150°C. The COD reagent also contains mercury that complexes chloride interference since chloride was present in the simulated SME products. Oxidizable organic compounds reduce the yellow colored dichromate ion ( $\text{Cr}_2\text{O}_7^{2-}$ ) to the green colored chromic ion ( $\text{Cr}^{3+}$ ). The amount of chromic ion produced is measured colorimetrically. The results are expressed as the number of milligrams of oxygen ( $\text{O}_2$ ) consumed per liter of sample (mg/L COD). This number can be converted to carbon equivalents on a molar basis since one mole of  $\text{O}_2$  and one mole of C are needed to form  $\text{CO}_2$ .

Initially a series of standards were analyzed ranging from 0  $\mu\text{g/mL}$  to 10,000  $\mu\text{g/mL}$  in order to determine linearity and the dynamic range of the COD method. The standards were prepared from a stock 10,000  $\mu\text{g/mL}$  COD standard purchased from Chemetrics. A graph and data results are in Figure 1 and Table 2. The method is very linear in this dynamic range with a correlation factor of 0.99997.

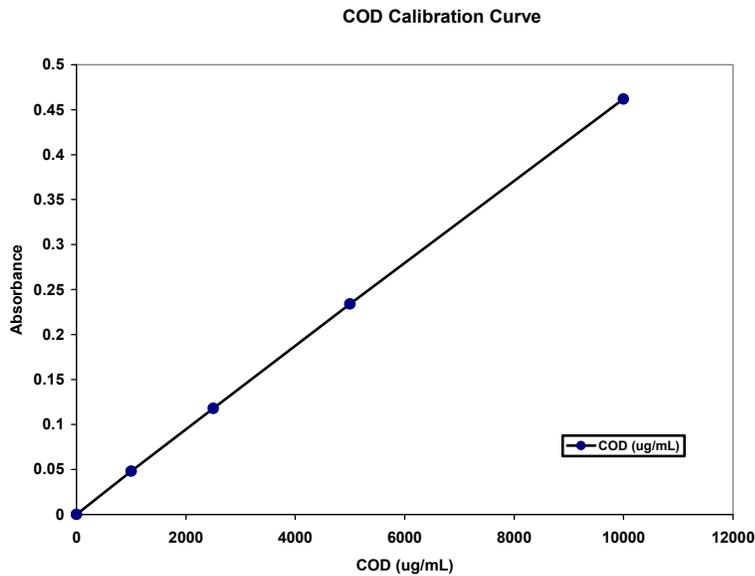


Figure 1. COD calibration curve.

**Table 2 COD Calibration Curve Data**

<u>Standard</u>	<u>Absorbance</u>
10000	0.462
5000	0.234
2500	0.118
1000	0.048
0	0

To further evaluate the COD technique and before beginning SB3 SME simulant analyses, a few initial experiments were done to determine the precision and effectiveness of the method on a few selected SB3 simulant samples. Two sets of SB3 simulant samples were analyzed in triplicate by the procedure recommended by the manufacturer. The manufacturer’s protocol for this procedure is to pipette a specific known volume of sample into the reagent vial. The vial is capped and mixed thoroughly and then placed in the digestion device. The sample is digested at 150°C for two hours, then removed to allow to cool for one hour before measuring the concentration of COD. The results of the two sets of SB3 SME simulant samples are in Table 3.

**Table 3 Reproducibility of SME COD Analyses**

<u>Samples</u>	<u>Conc.</u>	<u>AVE</u>	<u>SD</u>	<u>RSD</u>
SB3-3-1	17,611			
SB3-3-2	19,037			
SB3-3-3	18,418	18,355	715	3.90
SB3-1-25-202-1	10,619			
SB3-1-25-202-2	11,269			
SB3-1-25-202-3	10,986	10,958	326	2.97

Based on these results with relative standard deviations <5% it was decided to proceed with the analyses of the COD/SB3 SME simulants. All 200 SB3 simulant samples prepared were analyzed one time using the procedure recommended by the manufacturer.

Considering that some of these simulant samples contain particles of coal, it was felt that it would be useful to know the effectiveness of oxidation of the particles by the COD technique outside the simulant matrix. The particles were analyzed using the procedure recommended by the manufacturer. One to ten particles were added to the reagent vials

and allowed to digest for two hour at 150°C. It should be noted that in each case the coal was not totally consumed by the reagent and particles of coal remained in the vials after the two hour digestion. However, as one might expect, as the number of coal particles increased so did the COD concentration. The relative sizes and surface area of the coal particles was not measured as the goal of these scoping experiments was simply to see if more coal gave an increased COD response. The results are given in Table 4.

**Table 4 Increase in COD Response with Coal Concentration**

<u>Sample</u>	<u>Conc.</u>
Coal 1 particle	1010
Coal 3 particles	3370
Coal 5 particles	4310
Coal 10 particles	7400

Upon completion of the COD analyses of the supplied SB3 SME simulant sample series, the results were reviewed. There seemed to be more scatter in the data than was expected. It was decided to investigate the COD methodology more closely and determine if the technique could be modified to improve the precision and not complicate the procedure in the event it may eventually be used in DWPF.

### **3.2.2 Modified COD Analyses to Accommodate Coal Digestion**

After evaluation of the COD method and a series of experiments, two minor modifications were made in the procedure to better accommodate SB3 SME simulant oxalate/coal containing samples and improve the precision.

The two modifications made in the procedure were to:

- 1) remove  $\sim\frac{1}{4}$  inch of the pipet tip to ensure representative sampling of the high wt% solids SME products, e.g. open the sampling orifice and
- 2) lengthen the digestion time from two hours to four hours to enhance the digestion of organic material within the sample.

Table 5 demonstrates the effects of sample digestion time at two, four and six hours on COD response. Based on this data a four hour digestion appears to be the optimum length of time for the coal containing SB3 simulant samples.

**Table 5 Results of Various Digestion Times on COD Response**

<u>Sample ID</u>	<u>Conc (ug/mL)</u>	<u>Digested (hours)</u>
SB3-6(2)-25-202 (.0615 Coal)	10,333	2
SB3-6(4)-25-202 (.0615 Coal)	11,432	4
SB3-6(6)-25-202 (.0615 Coal)	11,156	6
SB3-7(2)-25-202 (.0615 Coal)	8,685	2
SB3-7(4)-25-202 (.0615 Coal)	9,870	4
SB3-7(4)-25-202 (.0615 Coal)	9,610	6

To validate the precision of the modified COD procedure a set of three coal containing SB3 simulant samples were analyzed in triplicate. The results are shown in the table below. The precision is good with relative standard deviations of <2.5% for all three samples. It should be noted before digestion the vials are cloudy (murky looking). After digestion the vials were clear and the only particles that seemed to be remaining were attributed to a small amount of frit at the bottom of the vial.

**Table 6 Precision of Modified COD Method to Accommodate Coal Digestion**

<u>Sample</u>	<u>COD (mg/L)</u>	<u>AVE</u>	<u>SD</u>	<u>RSD</u>
SB3-5(1)-35-202 (.041 Coal)	19,867			
SB3-5(2)-35-202 (.041 Coal)	19,362			
SB3-5(3)-35-202 (.041 Coal)	19,174	19,468	358	1.84
SB3-6(1)-35-202 (.041 Coal)	12,202			
SB3-6(2)-35-202 (.041 Coal)	12,746			
SB3-6(3)-35-202 (.041 Coal)	12,289	12,412	292	2.35
SB3-7(1)-35-202 (.041 Coal)	11,958			
SB3-7(2)-35-202 (.041 Coal)	11,472			
SB3-7(3)-35-202 (.041 Coal)	11,473	11,634	280	2.41

### 3.2.3 Effects of Nitric Acid on COD Analyses

The presence of nitric acid in the SME samples and its affect on the COD technique was of concern. A few experiments were performed to investigate this issue. Two solutions of sodium oxalate were prepared, one solution without nitric acid and the other in 0.28 M nitric acid. Both solutions were analyzed for COD twelve times. As shown in the table below, samples prepared in 0.28 M nitric acid gave a 3.30% higher COD value than those samples prepared without nitric acid and the precision for both solution remained at <2.5% RSD.

### 3.3 QUALITY ASSURANCE

All the data reported in this study were developed under the quality assurance given in technical task plan WSRC-RP-2002-00341 [21]. The research program and task plan were developed to address TTR - HLW/DWPF/TTR-02-0017. The data are recorded in notebooks WSRC-NB-2002-00156, WSRC-NB-2002-00199, WSRC-NB-2003-00034, and WSRC-NB-99-00100.

Table 7 HNO<sub>3</sub> Effects on COD Analyses

<u>Sample ID</u>	<u>Conc.</u>	<u>Average</u>	<u>SD.</u>	<u>RSD</u>
SO-13	1960			
SO-14	2040			
SO-15	1920			
SO-16	1960			
SO-17	1960			
SO-18	2040	1,990	38	1.89
SO-19	2000			
SO-20	2000			
SO-21	2040			
SO-22	1980			
SO-23	2000			
SO-24	1980			
SO/NA-13	1980			
SO/NA-14	2120			
SO/NA-15	2110			
SO/NA-16	2100			
SO/NA-17	2020			
SO/NA-18	2130	2,058	50	2.44
SO/NA-19	2040			
SO/NA-20	2050			
SO/NA-21	2010			
SO/NA-22	2090			
SO/NA-23	2020			
SO/NA-24	2020			
Blank-1	0			
Blank-2	0			
Blank-3	0			

Notes:

SO= Sodium Oxalate  
 SO/NA = Sodium Oxalate/Nitric Acid  
 NA Conc = 0.28 Molar  
 SO Conc. = 0.1504 Molar  
 SO/NA Conc. = 0.1496

## 4.0 REGRESSION ANALYSIS OF MODEL CRUCIBLE DATA

### 4.1 REPORTING UNITS

In this study, COD measurement and its usage for prediction of REDOX and control information pertinent to DWPF (i.e., feed or glass produced from simulated waste sludge) is examined. The COD analyses are reported in ppm (ug/mL or mg/L) and converted to mol/L by dividing by the atomic weight of O<sub>2</sub> times 1000. Alternatively the amount of O<sub>2</sub> can be converted to the amount of C as was done in the 1982 Soper reference [17] by assuming that 1 mole of carbon is equivalent to one mole of O<sub>2</sub> because it creates one mole of CO<sub>2</sub> (see Equation 4). The mol/L units need to be converted to mol/kg in order to be consistent with the units used in the development of the Electron Equivalents REDOX model at the given SME wt% solids using the SME density, as shown below.

**Equation 5**

$$COD(\text{mol / kg slurry}) = \frac{COD(\text{mol / L})}{\rho_1(\text{kg / L})}$$

*where  $\rho_1$  = density at initial total solids content  $T_1$*

For example, at 35 wt% solids ( $T_1$ ), the SME density is about 1.256 kg/L. This density value is from a correlation of density versus total solids that was derived from SRTC and DWPF data. This correlation is shown in Figure 2. The supporting data for this correlation is given in Reference 10. Because the frit does not have a COD response [17] and acts as diluent, the concentration in mol/kg is adjusted to the SRAT wt% solids as follows:

**Equation 6**

$$COD^{SRAT}(\text{mol / kg slurry}) = \frac{COD(\text{mol / kg slurry})}{\left( \frac{SRAT \text{ slurry}(gms)}{SRAT \text{ slurry}(gms) + FRIT (gms)} \right)} \times \text{wt\% SRAT solids}$$

This method normalizes the concentration on a ‘per mass of SRAT slurry’ basis. The COD<sup>SRAT</sup> (mol/kg slurry) values are given in Table 1.

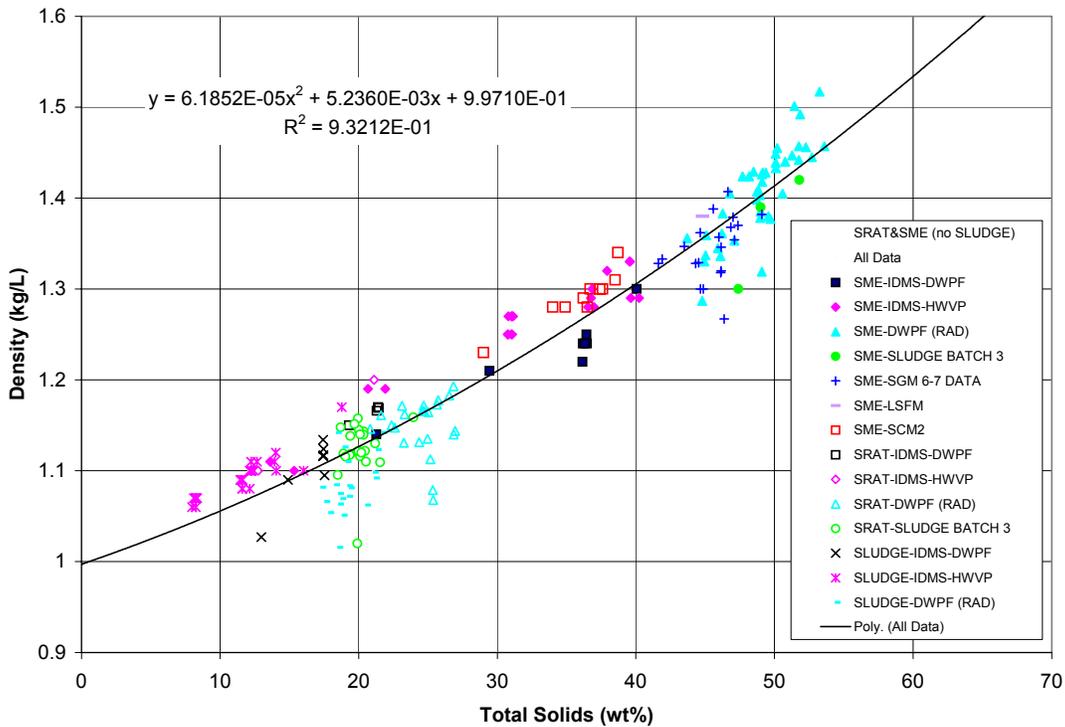


Figure 2. Density correlated to total solids.

## 4.2 COD MODELING

The independent variable (X) in the Molar Electron Equivalents ( $\xi$ ) REDOX model is a combination of concentrations as given in Equation 1, e.g.  $\xi$  (mol/kg feed) =  $2[F] + 4[C] + 4[O_T] - 5[N] - 2[Mn]$

where [F] = formate (mol/kg feed)  
 [C] = coal (carbon) (mol/kg feed)  
 [O<sub>T</sub>] = total oxalate, soluble and insoluble (mol/kg feed)  
 [N] = nitrate + nitrite (mol/kg feed)  
 [Mn] = manganese (mol/kg feed)

Only fifty three glasses of the ~200 fabricated were used during the development of the Electron Equivalents model (see discussion of measured REDOX acceptability and glass sample inhomogeneity in Reference 10). Of the 53 samples used to develop the REDOX correlation only 37 had COD analyses performed on them (see Table 1). The COD analyses reported are those measured before the COD procedure was modified to give a more accurate response for solid coal particles (see Section 3.3.2).

A fit of the REDOX data which includes the 53 glasses from Table 1 and 76 additional REDOX measurements from the 1997 study [16] to  $\xi$  is given in Figure 3 for comparison

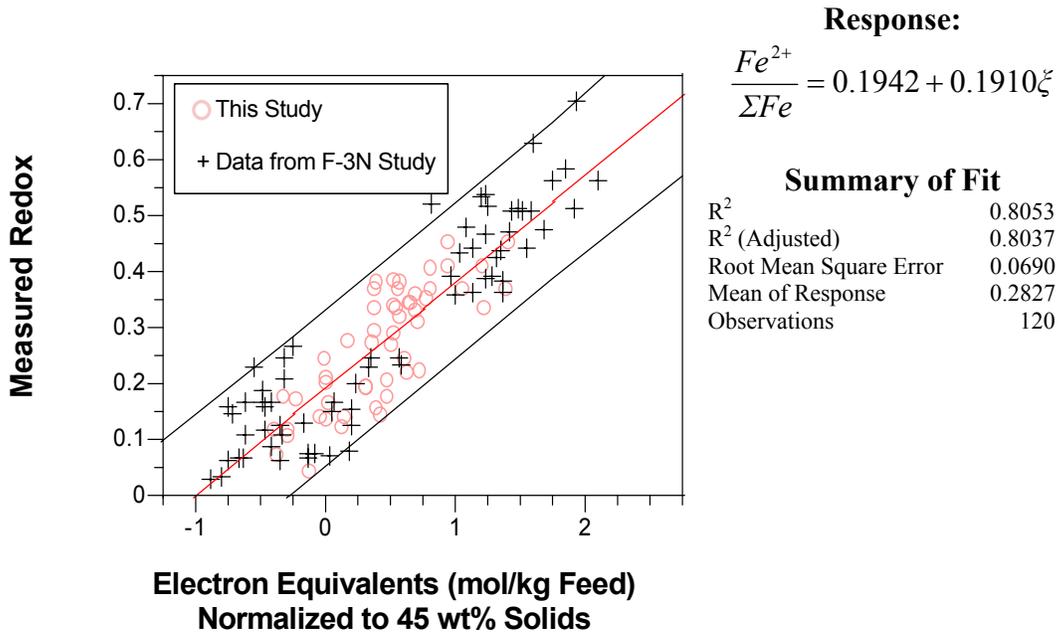


Figure 3. REDOX model with formate, oxalate, coal, nitrate, and manganese normalized for 45 wt% solids.

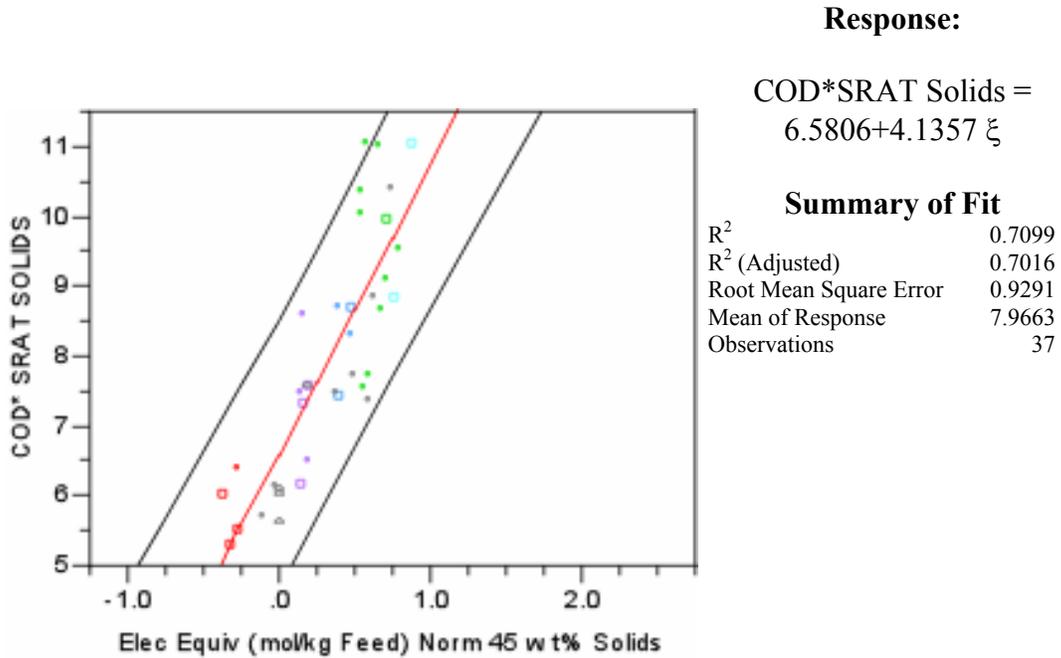


Figure 4. Comparison of the Electron Equivalents to the COD response weighted by the SRAT solids.

to the relationship between  $\xi$  and the COD response (Equation 6) in Table 1. Even though the  $\xi$  range of the COD response is narrower than the range of the REDOX response, a comparison of the two figures demonstrates that the COD response is similar to the Electron Equivalents response and monitors the relative strength of the various reductants and oxidants present in the feed.

However, when the COD response is modeled against the measured REDOX, the  $R^2$  of the correlation shown in Figure 5 is not very accurate, e.g.  $R^2=0.55$ . This is most likely due to the fact that the  $Fe^{+2}/\Sigma Fe$  range (0.04-0.45) over which the COD model is fit is almost half the  $Fe^{+2}/\Sigma Fe$  range (0.03-0.71) measured and used to develop the Electron Equivalents model.

Based on the similarity of the response of COD and the calculated electron equivalents further investigation of the COD methodology for DWPF may be warranted. The preliminary analyses presented here, based primarily on COD analyses that were measured before the procedure was modified, indicate that the COD analyses may be responding not only to the reductants in the feed but to the sum of the reductants and oxidants. This would eliminate the need for all but two measurements to predict feed REDOX (COD and SRAT wt% solids).

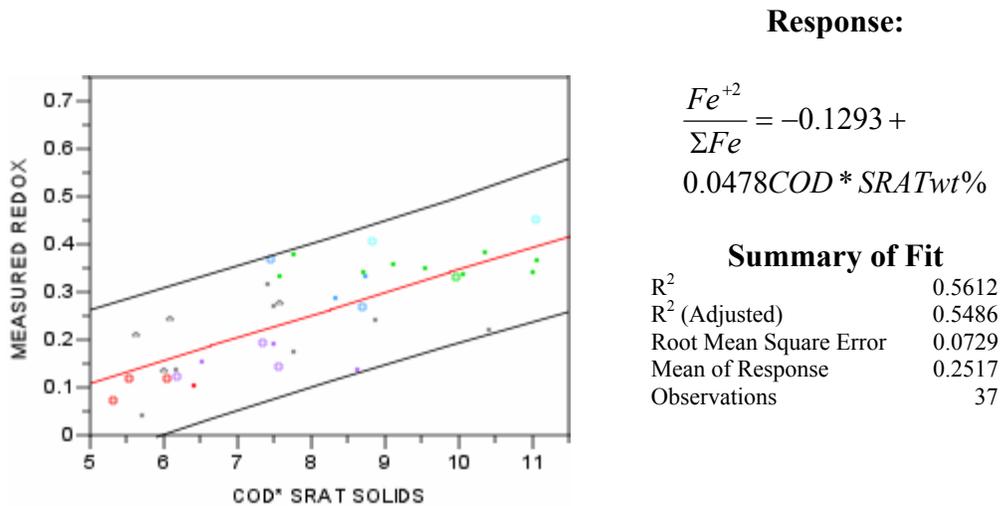


Figure 5. Comparison of the measured REDOX to the COD response weighted by the SRAT solids wt%.

## 5.0 CONCLUSIONS

- The COD response is similar to the Electron Equivalents ( $\xi$ ) response indicating that COD is capable of monitoring the multiple REDuction and OXidation (REDOX) electron transfers not previously thought possible
- The COD response (based on two analytic measurements) does not track the measured  $\text{Fe}^{+2}/\Sigma\text{Fe}$  as well as the Electron Equivalents model (based on nine analytic measurements)
- The poor COD REDOX correlation is likely due to the fact that the COD data modeled was measured before the accuracy of the COD technique was improved
- The poor COD REDOX correlation may also be because the  $\text{Fe}^{+2}/\Sigma\text{Fe}$  data range was half the range of the data modeled with the Electron Equivalents model, e.g. COD analyses were not available for the 1997 REDOX model data used to fit the Electron Equivalents model
- Based on the similarity of the response of COD and the calculated Electron Equivalents further investigation of the COD methodology for DWPF REDOX control may be warranted.
- The particle size of the coal varied throughout the study but the COD analyses were not perfected to the point that differences in the reactivity of the coal as a function of size could be assessed. These particle sizes are larger than those currently being analyzed in the SB3 sludge.

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